

Chem 331, Physical Chemistry I Final

Monday, December 14, 2007

200 points total

Definitions:	$H = U + PV$	$A = U - TS$
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Properties of Matter:	$C_v = \left(\frac{\partial U}{\partial T}\right)_V$	$C_P = \left(\frac{\partial H}{\partial T}\right)_P$
	$\beta = \frac{1}{V} \left(\frac{\partial V}{\partial T}\right)_P$	$\kappa = -\frac{1}{V} \left(\frac{\partial V}{\partial P}\right)_T$
		$\frac{\beta}{\kappa} = \left(\frac{\partial P}{\partial T}\right)_V$

Fundamental equations	Maxwell relations
$dU = TdS - PdV + \sum \mu_i dn_i$	$\left(\frac{\partial P}{\partial S}\right)_V = -\left(\frac{\partial T}{\partial V}\right)_S$
$dA = -SdT - PdV + \sum \mu_i dn_i$	$\left(\frac{\partial P}{\partial T}\right)_V = \left(\frac{\partial S}{\partial V}\right)_T$
$dH = TdS + VdP + \sum \mu_i dn_i$	$\left(\frac{\partial T}{\partial P}\right)_S = \left(\frac{\partial V}{\partial S}\right)_P$
$dG = -SdT + VdP + \sum \mu_i dn_i$	$\left(\frac{\partial S}{\partial P}\right)_T = -\left(\frac{\partial V}{\partial T}\right)_P$
$dS = \frac{dq_{rev}}{T}$	

Some derived relations:

$$\frac{dP}{dT} = \frac{\Delta \bar{H}}{T \Delta \bar{V}}, \quad \left(\frac{T_f}{T_i}\right) = \left(\frac{V_f}{V_i}\right)^{1-\frac{C_p}{C_v}}$$

$$\frac{d(\ln P)}{d(1/T)} = -\frac{\Delta \bar{H}}{R}, \quad C_p - C_v = nR$$

$$\frac{d \ln K}{dT} = \frac{\Delta H_r^\circ}{RT^2}, \quad \left(\frac{\partial G/T}{\partial T}\right)_P = -\frac{H}{T^2}$$

$$\Delta G_r = \Delta G_r^\circ + RT \ln Q$$

$$\gamma_i^R = \frac{P_i}{x_i P_i^*}, \quad \gamma_i^H = \frac{P_i}{x_i k_i}, \quad a_{x,i} = \gamma_i x_i, \quad a_{m,i} = \gamma_i m_i, \quad a_i = \frac{P_i}{P_i^*}, \quad f_i = \phi_i P_i; \quad v = v_+ + v_-$$

$$a(A_{v+}B_{v-}) = a_\pm^v = (\gamma_\pm m_\pm)^v; \quad \gamma_\pm = (\gamma_+^{v_+} \gamma_-^{v_-})^{1/v}, \quad m_\pm = (m_+^{v_+} m_-^{v_-})^{1/v} = (v_+^{v_+} v_-^{v_-})^{1/v} m$$

$$I = \frac{1}{2} \sum m_i z_i^2, \quad \ln \gamma_\pm = 1.173 z_+ z_- \sqrt{I}, \quad \mu_i = \mu_i^o + RT \ln a_i, \quad \sum n_i d\mu_i = 0$$

$$R = 8.3145 \text{ J K}^{-1} \text{ mol}^{-1} = 0.082058 \text{ L atm K}^{-1} \text{ mol}^{-1}; \quad 1 \text{ bar} = 10^5 \text{ Pa}$$