

KEY

Chem 534: Second Exam (100 points total)

Thursday, November 19, 2015

Show all work. Miscellaneous equations and other info is on the back page.

(1) (25 pts) Consider the reaction $\text{CO}_2(\text{g}) + \text{H}_2(\text{g}) \rightleftharpoons \text{CO}(\text{g}) + \text{H}_2\text{O}(\text{g})$ at 1 bar and 900 K, as well as the following data:

$$q_{\text{rot}}(\text{H}_2) = 5.1397 \quad q_{\text{rot}}(\text{CO}) = 323.90 \quad \tilde{B}(\text{CO}_2) = 0.3902 \text{ cm}^{-1}$$

$$\tilde{A}(\text{H}_2\text{O}) = 27.877 \text{ cm}^{-1} \quad \tilde{B}(\text{H}_2\text{O}) = 14.512 \text{ cm}^{-1} \quad \tilde{C}(\text{H}_2\text{O}) = 9.285 \text{ cm}^{-1}$$

$$\Delta E_r = \sum D_e(\text{CO}_2) + D_e(\text{H}_2) - D_e(\text{CO}) - \sum D_e(\text{H}_2\text{O}) = -125.1 \text{ kJ/mol}$$

this should have been -29.89 kJ/mol

(a) Calculate the rotational contribution to the equilibrium constant K_e .

$$\Theta_r(\text{CO}_2) = \frac{0.3902}{0.6952} = 0.5613 \text{ K} \quad q_{\text{rot}}(\text{CO}_2) = \frac{1}{2} \left(\frac{900}{0.5613} \right)$$

$$= 801.71$$

$$\Theta_A(\text{H}_2\text{O}) = \frac{27.877}{0.6952} = 40.099 \text{ K}$$

$$\Theta_B = 20.875 \text{ K} \quad \Theta_C = 13.356 \text{ K}$$

$$q_{\text{rot}}(\text{H}_2\text{O}) = \frac{\pi^{1/2}}{2} \left(\frac{900^3}{(40.099)(20.875)(13.356)} \right)^{1/2} = 226.30$$

$$K_{\text{rot}} = \frac{(323.90)(226.30)}{(801.71)(5.1397)} = 17.79$$

(b) Now compute the electronic contribution.

$$g_{\text{elec}} = g_{\text{el}} e^{\sum D_e/kT} = e^{\sum D_e/kT} \quad \text{since all singlets.}$$

$$K_{\text{elec}} = \frac{e^{\frac{D_e(\text{CO})}{kT}} e^{\sum D_e(\text{H}_2\text{O})/kT}}{e^{\sum D_e(\text{CO}_2)/kT} e^{D_e(\text{H}_2)/kT}}$$

$$= \frac{[D_e(\text{CO}) + \sum D_e(\text{H}_2\text{O}) - \sum D_e(\text{CO}_2) - D_e(\text{H}_2)]}{kT}$$

$$= e^{-\Delta E_r/RT} = e^{-125.1 \times 10^3 / (8.314)(900)} = 1.82 \times 10^7$$

(2) (25 pts) Consider again the ClO_2 radical in its 2B_1 electronic ground state.

(a) Starting from $S = kT \left(\frac{\partial \ln Q}{\partial T} \right)_{N,V} + k \ln Q$, derive the expression for the rotational contribution to the entropy and then calculate it for $T=300$ K (note: $\Theta_A = 2.489$ K, $\Theta_B = 0.481$ K, $\Theta_C = 0.403$ K)

$$Q = \frac{q^N}{N!} \quad \text{but } N! \text{ goes w/ translation.}$$

$$S_r = NkT \left(\frac{\partial \ln q_{\text{rot}}}{\partial T} \right)_V + Nk \ln q_{\text{rot}} \quad q_{\text{rot}} = \frac{\pi^{1/2}}{2} \left(\frac{T^3}{\Theta_A \Theta_B \Theta_C} \right)^{1/2}$$

$$\ln q_{\text{rot}} = \ln \left(\frac{\pi^{1/2}}{2} \right) + \frac{1}{2} \ln T^3 - \frac{1}{2} \ln (\Theta_A \Theta_B \Theta_C)$$

$$= \ln \left(\frac{\pi^{1/2}}{2} \right) + \frac{3}{2} \ln T - \frac{1}{2} \ln (\Theta_A \Theta_B \Theta_C)$$

$$S_r = \frac{3}{2} Nk + Nk \ln \left[\frac{\pi^{1/2}}{2} \left(\frac{T^3}{\Theta_A \Theta_B \Theta_C} \right)^{1/2} \right]$$

$$= (1.5)(8.314) + (8.314) \ln \left[\frac{\pi^{1/2}}{2} \left(\frac{300^3}{(2.489)(0.481)(0.403)} \right)^{1/2} \right]$$

$$= 85.6 \text{ J/mol}\cdot\text{K}$$

(b) Now do the analogous exercise as (a) for the electronic contribution.

$$g_{\text{elec}} = g_{\text{el}} e^{\frac{\sum D_e}{kT}} \quad \ln g_{\text{elec}} = \ln g_{\text{el}} + \frac{\sum D_e}{kT}$$

$$S_{\text{elec}} = NkT \left(-\frac{\sum D_e}{kT^2} \right) + Nk \left(\ln g_{\text{el}} + \frac{\sum D_e}{kT} \right)$$

$$= Nk \ln g_{\text{el}}$$

$$= (8.314) (\ln 2) = 5.763 \text{ J/mol}\cdot\text{K}$$

- (3) (15 pts) In the high temperature limit, calculate the constant volume molar heat capacity of NH_3 given that the vibrational contribution to $C_{v,m}$ is 2.21 J/mol-K (hint: first obtain an expression for the total energy not including the vibrational contribution)

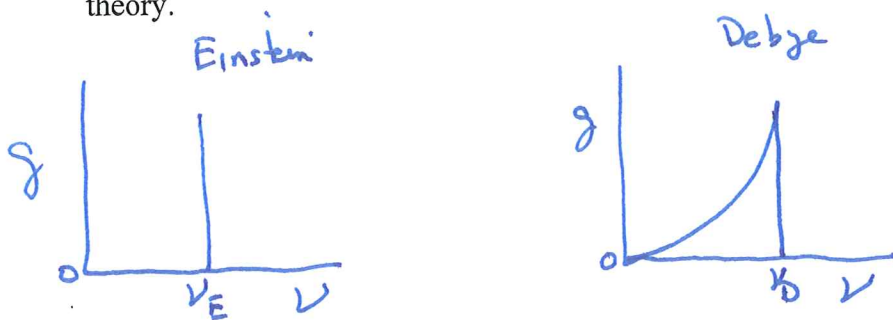
$$E = kT^2 \left(\frac{\partial \ln Q}{\partial T} \right) \quad \text{w/o vibrations} \quad E = \frac{3}{2} NkT \quad (\text{trans}) + \frac{3}{2} NkT \quad (\text{rot}) - N \sum E_e \quad (\text{elec})$$

$$C_v = \left(\frac{\partial E}{\partial T} \right)_v = 3Nk = 3R \quad \text{w/o vibrations}$$

$$C_v = (3)(8.314) + 2.21 = 27.15 \text{ J/mol}\cdot\text{K}$$

- (4) (20 pts.) Consider a perfect monatomic crystal

- (a) Sketch and compare the spectral density function $g(\nu)$ that is used in Einstein theory and the Debye theory.



Einstein: single frequency for all oscillators

Debye: treat acoustic phonons exactly, introduce a cutoff frequency ν_D

- (b) Briefly, what was the main drawback of the Einstein treatment? How did the Debye theory overcome these shortcomings? Include a short rationalization on why it works as well as it does. (continue on back if needed)

Einstein treatment led to incorrect behavior ^{of C_v} at low T , i.e., it did not yield a T^3 dependence. Debye treated the low frequency modes exactly since these are populated at low T .

(5) (15 pts.) Consider the classical 2-dimensional harmonic oscillator with Hamiltonian

$$H = \frac{p_x^2 + p_y^2}{2m} + \frac{1}{2}(k_x x^2 + k_y y^2)$$

(a) Predict the total energy of this system using the equipartition theorem.

$$E = 4 \left(\frac{1}{2} N k T \right) = 2 N k T$$

or $2 k T$ for a single molecule

(b) Set up the full expression for the ensemble average of the square of the total momentum,

$p^2 = (p_x^2 + p_y^2)$. Do not solve, but label carefully.

$$\langle p^2 \rangle = \frac{\int_{-\infty}^{\infty} \int_{-\infty}^{\infty} \int_{-\infty}^{\infty} \int_{-\infty}^{\infty} p^2 e^{-H/kT} dx dy dp_x dp_y}{\int_{-\infty}^{\infty} \int_{-\infty}^{\infty} \int_{-\infty}^{\infty} \int_{-\infty}^{\infty} e^{-H/kT} dx dy dp_x dp_y}$$

$$\int_{-\infty}^{\infty} \int_{-\infty}^{\infty} \int_{-\infty}^{\infty} \int_{-\infty}^{\infty} e^{-H/kT} dx dy dp_x dp_y$$

$$\text{where } p^2 = p_x^2 + p_y^2$$

$$H = \frac{p_x^2 + p_y^2}{2m} + \frac{1}{2}(k_x x^2 + k_y y^2)$$

Miscellaneous Information

$$k = 1.381 \times 10^{-23} \text{ J/K} = 0.6952 \text{ cm}^{-1}/\text{K}$$

$$N_A = 6.022 \times 10^{23} \text{ mol}^{-1} \quad R = 8.314 \text{ J mol}^{-1} \text{ K}^{-1}$$

$$h = 6.626 \times 10^{-34} \text{ J s}$$

$$c = 2.9979 \times 10^{10} \text{ cm/s}$$

$$1 \text{ atm} = 101,325 \text{ Pa} = 1.01325 \text{ bar}$$

$$E = kT^2 \left(\frac{\partial \ln Q}{\partial T} \right)_{V,N} \quad S = kT \left(\frac{\partial \ln Q}{\partial T} \right)_{N,V} + k \ln Q = k \ln Q + \frac{E}{T}$$

$$q_{\text{trans}} = \left(\frac{2\pi mkT}{h^2} \right)^{3/2} V$$

$$q_{\text{rot}} = \frac{\pi^{1/2}}{\sigma} \left(\frac{T^3}{\Theta_A \Theta_B \Theta_C} \right)^{1/2} \quad \text{or} \quad q_{\text{rot}} = \frac{1}{\sigma} \left(\frac{T}{\Theta_r} \right)$$

$$q_{\text{vib}} = \prod_{i=1}^{3n-6(5)} \frac{e^{-\theta_{v_i}/2T}}{1 - e^{-\theta_{v_i}/T}} \quad q_{\text{elec}} = \sum_i g_{ei} e^{-\epsilon_i/kT}$$

$$\Theta_r = \frac{hB}{k} = \frac{hc\tilde{B}}{k} \quad \Theta_r = \frac{hc\tilde{\nu}}{k}$$

$$\int_{-\infty}^{\infty} e^{-p^2/(2mkT)} dp = \sqrt{2\pi mkT}$$